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Reactions 113 SECTION

11.1 DESCRIBING
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321–329) This section

explains how to write
equations describing

chemical reactions using
appropriate symbols. It

also describes how to

write balanced chemical

equations when given

the names or formulas

of the reactants and

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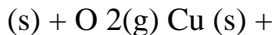
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the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions:

Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas.

Iron metal and chlorine gas react to form solid iron(III)chloride. CuS



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2 (page 324) 2. Sulfur

burns in oxygen to form

sulfur dioxide. Write a

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skeleton equation for
this chemical reaction.

Include appropriate
symbols from table

11.1. Analyze Step

1. Identify the relevant
concepts. Write the
formula for each
reactant and ...

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Reactions. In a chemical reaction, the reactants are written on the left and the products on the right. The arrow that separates them is called yield.

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113. SECTION 11.1

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REACTIONS (pages 321–329). This section

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17 (page 334). 17.

Complete the equations for these single-replacement reactions in aqueous solution.

Balance each equation.

Write “no reaction” if a

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this type must be to

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write the balanced chemical equation for the reaction. Inspection of Equation 11.4.1 shows that it is balanced as written, so we can proceed to the strategy outlined in Figure 11.4.1, adapting it as follows:

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