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~~View Answer. The equilibrium system~~

~~below shows the reaction of sour gas with~~

~~methane to produce carbon disulfide and~~

~~hydrogen gas. $\text{CH}_4(\text{g}) + 2\text{H}_2\text{S}(\text{g}) \rightarrow$~~

~~$\text{CS}_2(\text{g}) + 4\text{H}_2(\text{g})$; $\Delta H = +232.5 \text{ kJ...}$~~

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10⁻³ and 10³. Chemical equilibrium (definition) The state reached when the concentration of reactants and products remain constant over time. What is the general, reversible reaction for equilibrium? (With letters) $aA + bB \rightleftharpoons cC + dD$. Equilibrium equation (general with letters) $K_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}$.

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26 D5 Chemical equilibrium is said to be dynamic because A. the reaction proceeds quickly. B. the mass of the reactants is decreasing. C. the macroscopic properties are constant. D. both forward and reverse reactions are occurring. 27 D5

Equilibrium is a dynamic process because the A. macroscopic properties are not

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Study carefully the Sample Problem,
“ Using Le Châtelier ’ s Principle ” on
page 528 before you do questions 2.5 and
2.6 || **See your instructor to find out
what questions you should complete for
review. 2.3 (a) State in which direction the
equilibrium will shift for each of the
following, when there are

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This model should show the three states
that occur in the chemical equilibrium
process. As students work, walk around to
offer suggestions and guide. Have students:

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Answers the stress. 2. Answers may vary,
but should include changes in
concentration, volume, and temperature.
Some students may list pressure in place of

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Answers. 3. left; remove N_2 or H_2

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Chemical Equilibrium Constant: A

reversible chemical reaction system can

proceed in both directions from a set of

initial conditions (concentrations) towards

a state of chemical equilibrium.

~~What is the significance of the equilibrium~~

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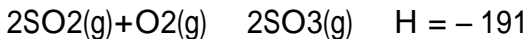
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The given reaction equation is:



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$2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$
 $\Delta H = -191 \text{ kJ}$. The equilibrium constant is a ratio of the forward and reverse ...

~~Consider the following equilibrium reaction. 2 ... - study.com~~

1) Give a balanced chemical equation which describes this equilibrium. 2) Explain what influence the removal of H_2 from the container will have on the equilibrium. 3) Explain what effect...

~~You are given a container in which $\text{H}_2(\text{g})$... - study.com~~

Solved: Consider the following equilibrium: $4\text{HCl}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g}) + 2\text{Cl}_2(\text{g})$; $K_c = 357$. An equilibrium mixture of the system was found to...

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Changing the temperature of a reaction at equilibrium alters both the equilibrium. (12) and the equilibrium position. When a reaction is. (13), which means it releases energy, lowering the temperature shifts the equilibrium to the (14) because the forward reaction liberates heat and removes the (15).

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